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Models of the Atom.

The Rutherford's model of the atom did not explain how an atom can emit light or the chemical properties of an atom. Plum Pudding Model Rutherford's Model. The Bohr Model.

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consists of an
element's symbol,
representing the
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that is surrounded by
dots, representing the
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electrons. <https://quizlet.com/6112461/che-chapter-5-electrons-in-atoms-flash-cards/> read more

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Although the speed of
all electromagnetic
waves in a vacuum is

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the same, waves can have different wavelengths and frequencies. As you can see from the equation on the previous page, wavelength and frequency are inversely related; in other words, as one quantity increases, the other decreases.

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CHE Chapter 5:
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consists of an element's symbol, representing the atomic nucleus and inner-level electrons, that is surrounded by dots, representing the atom's valence

electrons. Chemistry Chapter 5 Electrons In Atoms Test Answers Other Results for Chemistry Chapter 5 Test The Periodic Law Answers: Chemistry Chapter 5: The

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Atoms43 SECTION 5.1
MODELS OF THE ATOM
(pages 127-132) This
section summarizes
the development of
atomic theory. It also
explains the
significance of
quantized energies of
electrons as they relate

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to the quantum mechanical model of the atom. The Development of Atomic Models (pages 127-128) 1.

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Electrons occupy equal energy orbitals singly before beginning to pair up, because

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according to Hund's rule the most stable arrangement of electrons occurs when electrons with the same spin is as large as possible.

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istrom,Zike. Chapter 5
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Objects get their colors from reflecting only certain wavelengths when hit with white light. Light reflected from a green leaf is found to have a wavelength of $4.90 \times 10^{-7} \text{ m}$

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